

# Year 9: Chemical World- The Periodic Table

	Check	Date
<b>Revise assumed knowledge:</b> <b>SC4-16CW</b> describes the observed properties and behaviour of matter, using scientific models and theories about the motion and arrangement of particles  <b>SC4-17CW</b> explains how scientific understanding of, and discoveries about the properties of elements, compounds and mixtures relate to their uses in everyday life	<input type="checkbox"/>	
<b>CW2 The atomic structure and properties of elements are used to organise them in the Periodic Table. (ACSSU186)</b>		
<b>HISTORY OF THE PERIODIC TABLE</b>		
<b>5CW2a. identify the atom as the smallest unit of an element and that it can be represented by a symbol</b>	<input type="checkbox"/>	
<b>Define</b> the terms element and atom	<input type="checkbox"/>	
<b>Identify</b> the symbols for common elements	<input type="checkbox"/>	
<b>LITERACY SET 1: COSMOS ARTICLE</b>		
<b>Assessment: Oxford online test-</b> History of the Periodic table. Students to achieve 100% in Support and Consolidate <b>OR</b> Consolidate and Extend	<input type="checkbox"/>	
<b>ORGANISING THE ELEMENTS</b>		
<b>5CW2b. distinguish between the atoms of some common elements by comparing information about the numbers of protons, neutrons and electrons</b>	<input type="checkbox"/>	
<b>Define</b> the terms atomic number, atomic mass, proton, neutron and electron	<input type="checkbox"/>	
<b>Recall</b> key patterns within the periodic table (groups, periods, metals and non-metals, reactivity)	<input type="checkbox"/>	
<b>Identify</b> the location of protons, neutrons and electrons in an atom	<input type="checkbox"/>	
<b>Distinguish</b> between atomic and mass number	<input type="checkbox"/>	
<b>5CW2c. describe the organisation of elements in the Periodic Table using their atomic number</b>	<input type="checkbox"/>	
<b>Identify</b> metals, metalloids, non-metals and inert gases on a periodic table	<input type="checkbox"/>	
<b>Identify</b> and number the columns (Groups) and rows (Periods)	<input type="checkbox"/>	
<b>Describe</b> the organisation of common elements based on their atomic number	<input type="checkbox"/>	
<b>LITERACY SET 2: MIXED ACTIVITIES</b>		
<b>Assessment: Oxford online test-</b> Organising the elements Students to achieve 100% in Support and Consolidate <b>OR</b> Consolidate and Extend	<input type="checkbox"/>	
<b>ATOMIC STRUCTURE AND PROPERTIES</b>		
<b>5CW2d. relate the properties of some common elements to their position in the Periodic Table</b>	<input type="checkbox"/>	
<b>5CW2e. predict, using the Periodic Table, the properties of some common elements</b>	<input type="checkbox"/>	
<b>Describe</b> the general properties of metals	<input type="checkbox"/>	
<b>Describe</b> the properties of alkali (group I) metals	<input type="checkbox"/>	
<b>Describe</b> the properties of alkaline (group II) metals	<input type="checkbox"/>	

<b>Define and describe</b> the term transition metals	<input type="checkbox"/>	
<b>CODE: 9CW20 First-Hand investigation:</b> Reactivity of Metals (Oxford pg56)	<input type="checkbox"/>	
<b>Define and describe</b> the term non-metal	<input type="checkbox"/>	
<b>Define and describe</b> the term halogen (group VII)	<input type="checkbox"/>	
<b>Define and describe</b> the term noble gas (group VIII)	<input type="checkbox"/>	
<b>Define</b> the terms electron shells, valence electrons and ions	<input type="checkbox"/>	
<b>Describe</b> the concept of electron shells (2,8,8)	<input type="checkbox"/>	
<b>Explain</b> the link between periodic table groups and the number of valence electrons	<input type="checkbox"/>	
<b>Predict</b> the properties (conductivity, chemical reactivity, melting and boiling point) of some common elements based on their position in the periodic table	<input type="checkbox"/>	
<b>9CW21 First-Hand investigation:</b> Conductivity of ionic compounds (Oxford pg69)	<input type="checkbox"/>	
<i>5CW2f. outline some examples to show how creativity, logical reasoning and the scientific evidence available at the time, contributed to the development of the modern Periodic Table</i>	<input type="checkbox"/>	
<b>Recall</b> that the periodic table is organised by atomic number	<input type="checkbox"/>	
<b>Describe</b> the format of the current periodic table <u>compared</u> to previous versions	<input type="checkbox"/>	
<b>Describe</b> the contributions to the development of the periodic table by Alexandre-Emile Beguyer de Chancourtois, John Newlands, Julius Lothar Meyer, Dmitri Mendeleev and Henry Mosley	<input type="checkbox"/>	
<b>NUMERACY AND SKILLS SET</b>	<input type="checkbox"/>	
<b>Assessment: Oxford online test-</b> Atomic structure and properties Students to achieve 100% in Support and Consolidate <b>OR</b> Consolidate and Extend	<input type="checkbox"/>	
<b>Assessment:</b> Periodic Table Chapter Test	<input type="checkbox"/>	
<b>Comments and Suggested improvements</b>		
<p><b>Name:</b> _____ <b>Signature:</b> _____ <b>Date:</b> _____</p>		